

# Mononuclear Metal–O<sub>2</sub> Complexes Bearing Macrocyclic *N*-Tetramethylated Cyclam Ligands

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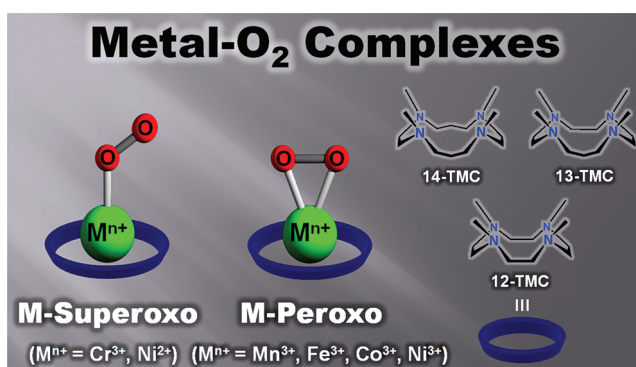
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## CONSPECTUS

**M**etalloenzymes activate dioxygen to carry out a variety of biological reactions, including the biotransformation of naturally occurring molecules, oxidative metabolism of xenobiotics, and oxidative phosphorylation. The dioxygen activation at the catalytic sites of the enzymes occurs through several steps, such as the binding of O<sub>2</sub> at a reduced metal center, the generation of metal–superoxo and –peroxo species, and the O–O bond cleavage of metal–hydroperoxo complexes to form high-valent metal-oxo oxidants. Because these mononuclear metal–dioxygen (M–O<sub>2</sub>) adducts are implicated as key intermediates in dioxygen activation reactions catalyzed by metalloenzymes, studies of the structural and spectroscopic properties and reactivities of synthetic biomimetic analogues of these species have aided our understanding of their biological chemistry. One particularly versatile class of biomimetic coordination complexes for studying dioxygen activation by metal complexes is M–O<sub>2</sub> complexes bearing the macrocyclic *N*-tetramethylated cyclam (TMC) ligand.

This Account describes the synthesis, structural and spectroscopic characterization, and reactivity studies of M–O<sub>2</sub> complexes bearing tetraazamacrocyclic *n*-TMC ligands, where M = Cr, Mn, Fe, Co, and Ni and *n* = 12, 13, and 14, based on recent results from our laboratory. We have used various spectroscopic techniques, including resonance Raman and X-ray absorption spectroscopy, and density functional theory (DFT) calculations to characterize several novel metal–O<sub>2</sub> complexes. Notably, X-ray crystal structures had shown that these complexes are end-on metal-superoxo and side-on metal-peroxo species. The metal ions and the ring size of the macrocyclic TMC ligands control the geometric and electronic structures of the metal–O<sub>2</sub> complexes, resulting in the end-on metal–superoxo versus side-on metal–peroxo structures. Reactivity studies performed with the isolated metal-superoxo complexes reveal that they can conduct electrophilic reactions such as oxygen atom transfer and C–H bond activation of organic substrates. The metal–peroxo complexes are active oxidants in nucleophilic reactions, such as aldehyde deformylation. We also demonstrate a complete intermolecular O<sub>2</sub>-transfer from metal(III)–peroxo complexes to a Mn(II) complex. The results presented in this Account show the significance of metal ions and supporting ligands in tuning the geometric and electronic structures and reactivities of the metal–O<sub>2</sub> intermediates that are relevant in biology and in biomimetic reactions.



## 1. Introduction

Metalloenzymes activate dioxygen to carry out a variety of biological reactions, including biotransformation of naturally occurring molecules, oxidative metabolism of xenobiotics, and oxidative phosphorylation.<sup>1</sup> The dioxygen activation at the active sites of the enzymes occurs through several steps, such as the binding of O<sub>2</sub> at a reduced metal

center, the generation of metal–superoxo and –peroxo species, and the O–O bond cleavage of metal–hydroperoxo complexes to form high-valent metal-oxo oxidants.<sup>1</sup> In heme and nonheme iron and copper enzymes, mononuclear metal–dioxygen adducts (M–O<sub>2</sub>), such as end-on and side-on metal–O<sub>2</sub> species (Scheme 1), are generated in the O<sub>2</sub> activation and play important roles as key intermediates in the

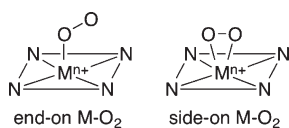
oxidation of organic substrates.<sup>1–4</sup> Such metal–O<sub>2</sub> intermediates have been investigated extensively in biomimetic studies to understand the structural and chemical properties of the intermediates that are short-lived and thus difficult to study in enzymatic reactions.<sup>5–11</sup> For example, the binding modes of the O<sub>2</sub> unit and the bond lengths of O–O and M–O bonds of synthetic metal–superoxo and –peroxo complexes were elucidated by X-ray crystallographic analysis. Spectroscopic characterization of the intermediates, such as O–O and M–O stretching frequencies determined in resonance Raman (rRaman) measurements,<sup>12</sup> supported the assignment of the binding modes of the O<sub>2</sub> unit in M–O<sub>2</sub> complexes. The reactivities of these M–O<sub>2</sub> complexes have been extensively investigated in electrophilic and nucleophilic oxidation reactions.<sup>13,14</sup> For instance, Fe(III)–O<sub>2</sub> complexes with heme and nonheme ligands have been synthesized as chemical models of cytochrome P450 aromatases and Rieske dioxygenases, respectively, and have shown reactivities in nucleophilic reactions such as aldehyde deformylation.<sup>4,14</sup> As biomimetic compounds of copper-containing enzymes, peptidylglycine- $\alpha$ -hydroxylating monooxygenase (PHM) and dopamine- $\beta$ -monooxygenase (D $\beta$ M), mononuclear Cu–O<sub>2</sub> complexes have been synthesized and characterized and their reactivities have been investigated in electrophilic reactions.<sup>7,8,10,15–17</sup> Fe(III)–superoxo species have been proposed as active oxidants in C–H bond activation of substrates by nonheme iron enzymes,

such as isopenicillin *N*-synthase and *myo*-inositol oxygenase.<sup>2</sup> Biomimetic studies provided indirect evidence that mononuclear nonheme Fe(III)–superoxo complexes are capable of activating C–H bonds of organic substrates, although such Fe(III)–superoxo species have yet to be isolated and characterized spectroscopically.<sup>18</sup>

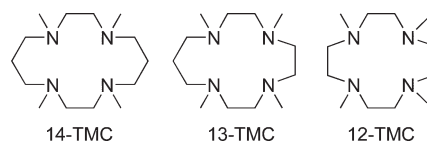
In addition to the iron and copper enzymes, Mn–O<sub>2</sub> complexes have been invoked as reactive intermediates in manganese-containing enzymes, such as in the oxygen evolving complex of photosystem II, manganese superoxide dismutase, and catalase.<sup>19</sup> In biomimetic studies, the first crystal structure of a side-on Mn(III)–peroxo complex, [Mn(TPP)(O<sub>2</sub>)]<sup>–</sup> (TPP = *meso*-tetraphenylporphyrin), was reported by Valentine and co-workers,<sup>20</sup> followed by the crystal structure of a nonheme side-on Mn(III)–peroxo complex by Kitajima and co-workers.<sup>21</sup> The chemistry of other M–O<sub>2</sub> complexes (M = Ti, V, Cr, Co, Ni, and the second and third row transition metals) has been well documented.<sup>13,22,23</sup> One representative example is the Co–O<sub>2</sub> complexes bearing salen, porphyrin, and tetraazamacrocyclic ligands as models of dioxygen-carrying proteins, such as hemoglobin and myoglobin, and as active oxidants in the oxidation of organic substrates.<sup>22,24,25</sup>

It has been demonstrated recently that *N*-tetramethylated cyclams (TMC),<sup>26</sup> such as *n*-TMC (*n* = 14 for 1,4,8,11-tetramethyl-1,4,8,11-tetraazacyclotetradecane, 13

**SCHEME 1.** End-on and Side-on Metal–O<sub>2</sub> Complexes



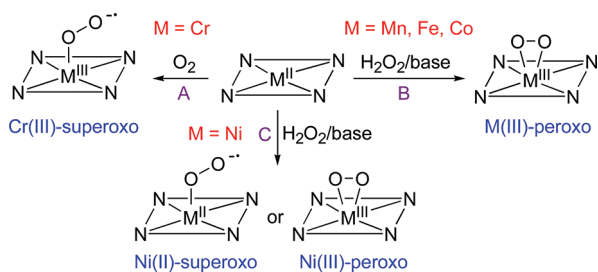
**SCHEME 2.** TMC Ligands Used in the Synthesis of M–O<sub>2</sub> Complexes



**TABLE 1.** Structures of Metal–O<sub>2</sub> Complexes Bearing Macrocyclic *n*-TMC Ligands<sup>a</sup>

	Chromium	Manganese	Iron	Cobalt	Nickel
14-TMC	<b>1</b>  Cr(III)-superoxo	<b>2</b>  Mn(III)-peroxo	<b>4</b>  Fe(III)-peroxo	<b>5</b>  Co(III)-peroxo	<b>8</b>  Ni(II)-superoxo
13-TMC		<b>3</b>  Mn(III)-peroxo		<b>6</b>  Co(III)-peroxo	
12-TMC		<b>Mn(III)-peroxo<sup>b</sup></b>  Mn(III)-peroxo <sup>b</sup>		<b>7</b>  Co(III)-peroxo	<b>9</b>  Ni(III)-peroxo

<sup>a</sup>There are no high-resolution crystal structures available for [Co<sup>III</sup>(14-TMC)(O<sub>2</sub>)]<sup>+</sup> (**5**) and [Ni<sup>II</sup>(14-TMC)(O<sub>2</sub>)]<sup>+</sup> (**8**). <sup>b</sup>The details of the high-resolution crystal structure of [Mn<sup>III</sup>(12-TMC)(O<sub>2</sub>)]<sup>+</sup> will be reported elsewhere.

**SCHEME 3.** Reactions Showing the Synthesis of Metal–Superoxo and –Peroxo Complexes

for 1,4,7,10-tetramethyl-1,4,7,10-tetraazacyclotridecane, and 12 for 1,4,7,10-tetramethyl-1,4,7,10-tetraazacyclodecane; see Scheme 2 for ligand structures), are versatile ligands in the biomimetic chemistry of dioxygen activation by metal complexes. A variety of metal–TMC complexes of oxo, superoxo, and peroxo ligands have been synthesized and characterized spectroscopically and/or structurally.<sup>27–40</sup> In particular, the first high-resolution crystal structures of Fe(IV)–oxo and Fe(III)–peroxo complexes with the 14-TMC ligand, such as  $[\text{Fe}^{\text{IV}}(14\text{-TMC}(\text{O}))_2]^{2+}$  and  $[\text{Fe}^{\text{III}}(14\text{-TMC}(\text{O}_2))]^+$ ,<sup>31,38</sup> were successfully obtained in nonheme iron models. In these studies, it has been shown that the ring size of the TMC ligands is an important factor that controls the geometric and electronic structures of the M–O<sub>2</sub> adducts. Such a ring size effect has been highlighted by two Ni–O<sub>2</sub> complexes bearing 14-TMC and 12-TMC ligands;<sup>34–36</sup>  $[\text{Ni}(14\text{-TMC}(\text{O}_2))]^+$  is an end-on Ni(II)-superoxo species,<sup>34</sup> whereas  $[\text{Ni}(12\text{-TMC}(\text{O}_2))]^+$  is a side-on Ni(III)-peroxo species.<sup>35</sup> Thus, significant progress has been achieved in understanding the chemical and physical properties of the biologically important metal–oxygen intermediates, such as high-valent metal–oxo and metal–O<sub>2</sub> complexes, by synthesizing their biomimetic compounds using the versatile tetraazamacrocyclic ligands, *n*-TMC. In this Account, we describe our recent results on the synthesis, structural and spectroscopic characterization, and reactivity studies of the M–O<sub>2</sub> complexes bearing *n*-TMC ligands (M = Cr, Mn, Fe, Co, Ni; *n* = 12, 13, 14) (Table 1). The effects of the metal ions and macrocyclic ring size on the structures and reactivities of the biologically important M–O<sub>2</sub> intermediates are discussed as well.

## 2. Synthesis and Characterization

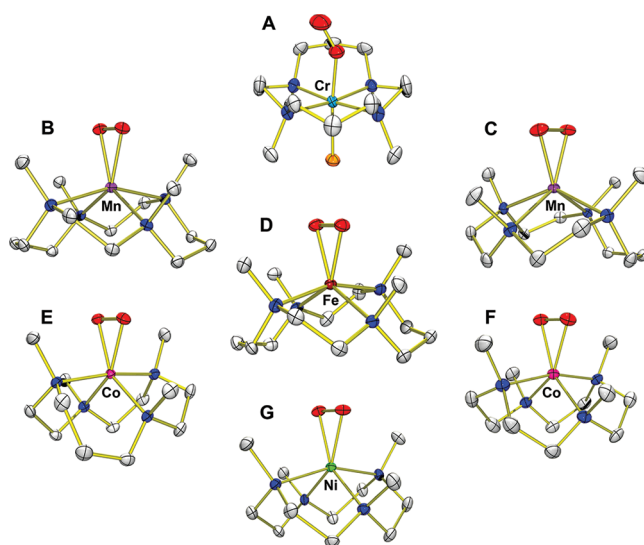
Metal–O<sub>2</sub> complexes were synthesized by reacting metal(II) complexes,  $[\text{M}(n\text{-TMC})]^{2+}$  (M = Cr, Mn, Fe, Co, Ni; *n* = 12, 13, 14), with O<sub>2</sub> or H<sub>2</sub>O<sub>2</sub> in the presence of base (e.g., triethylamine (TEA) or tetramethylammonium hydroxide (TMAH)), as summarized in Scheme 3. First, bubbling of O<sub>2</sub> to the solution of  $[\text{Cr}^{\text{II}}(14\text{-TMC}(\text{Cl}))]^+$  in CH<sub>3</sub>CN afforded the color change

**TABLE 2.** UV–Vis, ESI-MS, rRaman, EPR, Spin State, and X-ray Data of M–O<sub>2</sub> Complexes

entry	M–O <sub>2</sub> complex	UV–vis $\lambda$ , [nm] ( $\epsilon$ [M <sup>−1</sup> cm <sup>−1</sup> ])	ESI-MS $m/z$ <sup>16</sup> O <sub>2</sub> ( $m/z$ <sup>18</sup> O <sub>2</sub> )	rRaman [cm <sup>−1</sup> ] $\nu(\text{O}=\text{O})$ ( $\nu(\text{O}=\text{O})$ )	EPR	spin state	bond lengths M–O/O–O [Å]	ref
1	$[\text{Cr}^{\text{II}}(14\text{-TMC}(\text{O}_2)(\text{Cl}))]^+$ ( <b>1</b> )	331 (3800), 391 (290), 469 (150), 549 (240), 643 (130), 675 (140)	375.0 (379.0)	1170 (1104)	silent	S = 1	1.876(4)/1.231(6)	27
2	$[\text{Mn}^{\text{III}}(14\text{-TMC}(\text{O}_2))]^+$ ( <b>2</b> )	453 (490), 630 (120)	343.1 (347.1)		silent	S = 2	1.884(2)/1.403(4)	28
3	$[\text{Mn}^{\text{III}}(13\text{-TMC}(\text{O}_2))]^+$ ( <b>3</b> )	288 (2400), 452 (390), 615 (190)	329.1 (333.1)		silent	S = 2	1.859/1.410(4)	29
4	$[\text{Fe}^{\text{III}}(14\text{-TMC}(\text{O}_2))]^+$ ( <b>4</b> )	750 (600)	344.1 (348.1)	825 (781)	$g = 8.8, 5.9, 4.3$	S = 5/2	1.910/1.463(6)	31
5	$[\text{Co}^{\text{III}}(14\text{-TMC}(\text{O}_2))]^+$ ( <b>5</b> )	436 (250), 575 (100), 801 (90)	347.1 (351.1)		silent	S = 1		32
6	$[\text{Co}^{\text{III}}(13\text{-TMC}(\text{O}_2))]^+$ ( <b>6</b> )	348 (620), 562 (210), ~500 (170), ~710 (100)	333.1 (337.1)	902 (846)	silent	S = 0	1.855/1.438(4)	33
7	$[\text{Co}^{\text{III}}(12\text{-TMC}(\text{O}_2))]^+$ ( <b>7</b> )	350 (450), 560 (180), ~500 (150), ~710 (90)	319.1 (323.1)	902 (845)	silent	S = 0	1.866/1.4389(17)	33
8	$[\text{Ni}^{\text{II}}(14\text{-TMC}(\text{O}_2))]^+$ ( <b>8</b> )	336 (610), 414 (130), 690 (50)	346.1 (350.0)	1131 (1067)	$g = 2.29, 2.21, 2.09$	S = 1/2		34
9	$[\text{Ni}^{\text{III}}(12\text{-TMC}(\text{O}_2))]^+$ ( <b>9</b> )	345 (300), 400 (250), 646 (100), 899 (70)	318.0 (322.0)	1002 (945)	$g = 2.22, 2.17, 2.06$	S = 1/2	1.889/1.386(4)	35

from blue to violet (Scheme 3, route A).<sup>27</sup> The violet intermediate, [Cr<sup>III</sup>(14-TMC)(O<sub>2</sub>)(Cl)]<sup>+</sup> (**1**), was characterized with various spectroscopic methods (see Table 2, entry 1 for spectroscopic data). The superoxo character of the O<sub>2</sub> ligand was unambiguously assigned by rRaman spectroscopic and X-ray crystallographic analyses. In rRaman measurements, **1** prepared with <sup>16</sup>O<sub>2</sub> exhibits an isotopically sensitive band at 1170 cm<sup>-1</sup>, which shifts to 1104 cm<sup>-1</sup> in samples of **1** prepared with <sup>18</sup>O<sub>2</sub> (Table 2). The single crystal structure of **1** revealed the mononuclear end-on Cr–superoxo complex in a distorted octahedral geometry (Figure 1A). The O–O bond length (1.231 Å) of **1** is shorter than that of the side-on Cr(III)–superoxo complex (1.327 Å).<sup>41</sup> The relatively short O–O bond distance is consistent with the high frequency of the O–O stretching vibration observed in the rRaman measurements. All *N*-methyl groups of the 14-TMC ligand are oriented *anti* to the superoxo ligand, while the *syn* site is occupied by a Cl ligand.

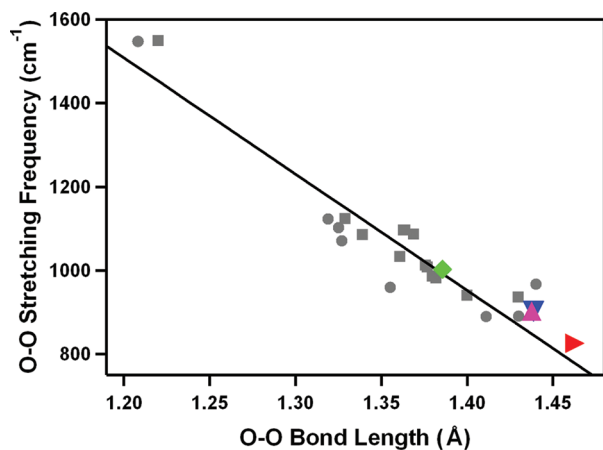
Different from the Cr–O<sub>2</sub> complex, other metal(II) complexes, such as Mn(II), Fe(II), Co(II), and Ni(II), did not react with O<sub>2</sub> to form M–O<sub>2</sub> complexes. Alternatively, metal–O<sub>2</sub> complexes were generated by reacting metal(II) complexes with H<sub>2</sub>O<sub>2</sub> in the presence of base (Scheme 3, routes B and C). First, [Mn<sup>III</sup>(14-TMC)(O<sub>2</sub>)]<sup>+</sup> (**2**) was synthesized by reacting [Mn<sup>II</sup>(14-TMC)]<sup>2+</sup> with H<sub>2</sub>O<sub>2</sub> in the presence of TEA in CH<sub>3</sub>CN (Scheme 3, route B; Table 2, entry 2 for spectroscopic data).<sup>28</sup> The ESI-MS of the green intermediate **2** exhibits a prominent ion peak at a mass-to-charge ratio (*m/z*) of 343.1, whose mass and isotope distribution pattern correspond to [Mn(14-TMC)(O<sub>2</sub>)]<sup>+</sup>. When **2** was synthesized using isotopically labeled H<sub>2</sub><sup>18</sup>O<sub>2</sub>, a mass peak corresponding to [Mn(14-TMC)(<sup>18</sup>O<sub>2</sub>)]<sup>+</sup> appeared at *m/z* 347.1. The four mass unit increase upon the substitution of <sup>16</sup>O with <sup>18</sup>O indicates the binding of an O<sub>2</sub> unit in **2**. The X-ray crystal structure of **2** revealed the mononuclear side-on Mn(III)–peroxo complex in a distorted octahedral geometry arising from the triangular MnO<sub>2</sub> moiety (Figure 1B). The O–O bond length of 1.403(4) Å is typical for a peroxo ligand bound to a transition metal ion, but is slightly shorter than those found in [Mn(TPP)(O<sub>2</sub>)]<sup>-</sup> (1.43 Å)<sup>20</sup> and [Mn(Tp<sup>P</sup>)(3,5-Pr<sub>2</sub>pZH)(O<sub>2</sub>)] (1.421 Å).<sup>21</sup> All *N*-methyl groups of the 14-TMC ligand are oriented *syn* to the peroxo ligand residing in the *N*-methyl pocket. Similarly, [Mn<sup>III</sup>(13-TMC)(O<sub>2</sub>)]<sup>+</sup> (**3**) was synthesized by reacting [Mn<sup>II</sup>(13-TMC)]<sup>2+</sup> with H<sub>2</sub>O<sub>2</sub> in the presence of TEA in CH<sub>3</sub>CN (Scheme 3, route B; Table 2, entry 3 for spectroscopic data).<sup>29</sup> The X-ray crystal structure of the green intermediate **3** revealed a monomeric six-coordinate Mn<sup>III</sup> cation with a



**FIGURE 1.** Crystal structures of M–O<sub>2</sub> complexes: (A) [Cr<sup>III</sup>(14-TMC)(O<sub>2</sub>)(Cl)]<sup>+</sup> (**1**); (B) [Mn<sup>III</sup>(14-TMC)(O<sub>2</sub>)]<sup>+</sup> (**2**); (C) [Mn<sup>III</sup>(13-TMC)(O<sub>2</sub>)]<sup>+</sup> (**3**); (D) [Fe<sup>III</sup>(14-TMC)(O<sub>2</sub>)]<sup>+</sup> (**4**); (E) [Co<sup>III</sup>(13-TMC)(O<sub>2</sub>)]<sup>+</sup> (**6**); (F) [Co<sup>III</sup>(12-TMC)(O<sub>2</sub>)]<sup>+</sup> (**7**); (G) [Ni<sup>III</sup>(12-TMC)(O<sub>2</sub>)]<sup>+</sup> (**9**).

side-on peroxo ligand and the O–O bond length of 1.403(4) Å (Figure 1C).

The Fe(III)–peroxo complex, [Fe<sup>III</sup>(14-TMC)(O<sub>2</sub>)]<sup>+</sup> (**4**), was prepared by reacting [Fe<sup>II</sup>(14-TMC)]<sup>2+</sup> with H<sub>2</sub>O<sub>2</sub> in the presence of TEA in CF<sub>3</sub>CH<sub>2</sub>OH (Scheme 3, route B).<sup>30,31</sup> The intermediate persisted for several hours at 0 °C, and the greater thermal stability of **4** in alcoholic solvents allowed us to isolate crystals.<sup>31</sup> The isolated intermediate **4** was fully characterized with various spectroscopic techniques (Table 2, entry 4), including UV–vis, ESI-MS, EPR, rRaman, and X-ray absorption spectroscopy (XAS). The rRaman spectrum of **4** shows two peaks that shift upon <sup>18</sup>O-substitution. The peak at 487 cm<sup>-1</sup> shifts to 468 cm<sup>-1</sup> with a <sup>16,18</sup>Δ value of 19 cm<sup>-1</sup> (<sup>16,18</sup>Δ (calcd) = 21 cm<sup>-1</sup>) and is assigned as an Fe–O stretch. The other isotopically sensitive peak in the rRaman spectrum appears at 825 cm<sup>-1</sup>, which shifts to 781 cm<sup>-1</sup> in the <sup>18</sup>O-labeled **4**; this feature is assigned as the O–O stretch with the <sup>16,18</sup>Δ value of 44 cm<sup>-1</sup> (<sup>16,18</sup>Δ (calcd) = 47 cm<sup>-1</sup>). These values are comparable to the Fe–O and O–O stretches in spectroscopically characterized high-spin side-on Fe(III)–peroxo complexes in heme and non-heme iron systems.<sup>31,42</sup> The X-ray crystal structure of **4** revealed the mononuclear side-on Fe(III)–peroxo complex in a distorted octahedral geometry (Figure 1D). The FeO<sub>2</sub> geometry is similar to the crystallographically characterized structure of naphthalene dioxygenase (NDO), where dioxygen binds in a side-on fashion at the mononuclear iron center (1.75 Å resolution, *r*<sub>O–O</sub> ca. 1.45 Å).<sup>4</sup> The structurally



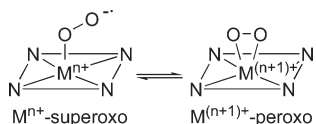
**FIGURE 2.** Plot of O–O stretching frequency ( $\text{cm}^{-1}$ ) versus O–O bond distance ( $\text{\AA}$ ) for side-on metal–O<sub>2</sub> complexes. Circles represent experimental data points, and squares represent theoretical ones.<sup>12</sup> The solid line represents a least-squares linear fit of the experimental and theoretical data. Data points for  $[\text{Fe}^{\text{IV}}(14\text{-TMC})(\text{O}_2)]^+$  (**4**) (red right triangle),  $[\text{Co}^{\text{III}}(13\text{-TMC})(\text{O}_2)]^+$  (**6**) (magenta up triangle),  $[\text{Co}^{\text{III}}(12\text{-TMC})(\text{O}_2)]^+$  (**7**) (blue down triangle), and  $[\text{Ni}^{\text{III}}(12\text{-TMC})(\text{O}_2)]^+$  (**9**) (green tilted square) are included in the diagram.

determined O–O bond length of 1.463(6)  $\text{\AA}$  is indicative of peroxo character of the O<sub>2</sub> group,<sup>12</sup> as assigned by rRaman experiments (vide supra). All four *N*-methyl groups of the 14-TMC ligand are oriented *syn* to the peroxo ligand. In the case of an Fe(IV)–oxo complex bearing the same 14-TMC ligand, the four *N*-methyl groups of  $[\text{Fe}^{\text{IV}}(14\text{-TMC})(\text{O})(\text{CH}_3\text{CN})]^{2+}$  are *anti* to the oxo ligand,<sup>38</sup> whereas those in a Sc<sup>3+</sup>-bound  $[\text{Fe}^{\text{IV}}(14\text{-TMC})(\text{O})]^{2+}$  are *syn* to the oxo ligand.<sup>39</sup> In addition, a good correlation was observed when the O–O bond length and O–O stretching frequency of **4** was fitted in the plot of O–O stretching frequency versus O–O bond length for side-on metal–O<sub>2</sub> complexes (Figure 2).<sup>12</sup>

The reaction of  $[\text{Co}^{\text{II}}(14\text{-TMC})(\text{CH}_3\text{CN})]^{2+}$  with H<sub>2</sub>O<sub>2</sub> in the presence of TEA in CH<sub>3</sub>CN afforded a gradual color change from pale purple to green (Scheme 3, route B).<sup>33</sup> The green intermediate,  $[\text{Co}^{\text{III}}(14\text{-TMC})(\text{O}_2)]^+$  (**5**), was assigned by ESI-MS and EPR experiments (Table 2, entry 5), and the structural information was obtained by XAS and density functional theory (DFT) calculations (vide infra). The Co(III)–peroxo complexes bearing 13-TMC and 12-TMC were synthesized as follows: Addition of H<sub>2</sub>O<sub>2</sub> to  $[\text{Co}^{\text{II}}(13\text{-TMC})(\text{CH}_3\text{CN})]^{2+}$  and  $[\text{Co}^{\text{II}}(12\text{-TMC})(\text{CH}_3\text{CN})]^{2+}$  in the presence of TEA in CH<sub>3</sub>CN afforded purple intermediates  $[\text{Co}^{\text{III}}(13\text{-TMC})(\text{O}_2)]^+$  (**6**) and  $[\text{Co}^{\text{III}}(12\text{-TMC})(\text{O}_2)]^+$  (**7**), respectively.<sup>33</sup> The spectroscopic (Table 2, entries 6 and 7) and structural (Figure 1E and F) characterization revealed that the peroxo ligand is bound to the cobalt ion in a side-on fashion. The structural properties of the CoO<sub>2</sub> core in **6** and **7** are nearly identical; the X-ray

structures of **6** and **7** revealed a mononuclear side-on Co(III)–peroxo complex in a distorted octahedral geometry arising from the triangular CoO<sub>2</sub> moiety, and the O–O bond distances of **6** and **7** were 1.438(6) and 1.4389(17)  $\text{\AA}$ , respectively (Table 2). All four *N*-methyl groups of the 13-TMC and 12-TMC ligands are oriented *syn* to the peroxo ligand in both complexes. The O–O bond stretching frequency of **6** and **7** was determined to be 902  $\text{cm}^{-1}$  by rRaman measurements, and we have observed a good correlation between the O–O stretching frequencies and the O–O bond lengths in **6** and **7** (Figure 2).<sup>12</sup>

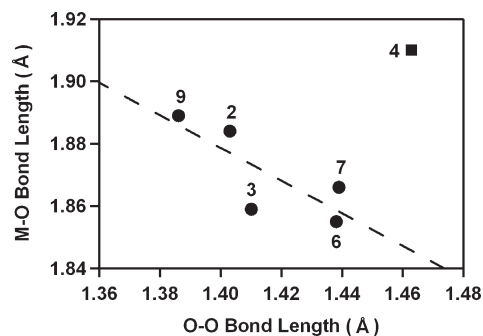
Different from the metal(III)–peroxo complex formation discussed above, a Ni(II)–superoxo complex,  $[\text{Ni}^{\text{II}}(14\text{-TMC})(\text{O}_2)]^+$  (**8**), was formed in the reaction of  $[\text{Ni}^{\text{II}}(14\text{-TMC})(\text{CH}_3\text{CN})]^{2+}$  and H<sub>2</sub>O<sub>2</sub> in the presence of base (TEA or TMAH) in CH<sub>3</sub>CN (Scheme 3, route C).<sup>34</sup> The formation of **8** was also observed when a low valent  $[\text{Ni}^{\text{I}}(14\text{-TMC})]^+$  complex was reacted with O<sub>2</sub>.<sup>34</sup> The formation of **8** was confirmed by various spectroscopic techniques (Table 2, entry 8). The rRaman data show that **8** exhibits an isotopically sensitive band at 1131  $\text{cm}^{-1}$ , which shifts to 1067  $\text{cm}^{-1}$  in samples of **8** prepared with <sup>18</sup>O-labeled H<sub>2</sub><sup>18</sup>O<sub>2</sub>. The O–O stretching frequency of **8** (1131  $\text{cm}^{-1}$ ) is comparable to that of **1** (1170  $\text{cm}^{-1}$ ), indicating the superoxo character of the O<sub>2</sub> unit in **8**. The structural information was obtained by XAS and DFT calculations (vide infra). Interestingly, a Ni(III)–peroxo complex,  $[\text{Ni}^{\text{III}}(12\text{-TMC})(\text{O}_2)]^+$  (**9**), was obtained by changing the supporting ligand from 14-TMC to 12-TMC.<sup>35</sup> The reaction of  $[\text{Ni}^{\text{II}}(12\text{-TMC})(\text{CH}_3\text{CN})]^{2+}$  with H<sub>2</sub>O<sub>2</sub> in the presence of TEA in CH<sub>3</sub>CN produces the green intermediate, **9**, which exhibits characteristic absorption bands that are distinct from those of **8** (Scheme 3, route C; Table 2, entry 9 for spectroscopic data). The rRaman spectra of **9** show the peroxo character of the O<sub>2</sub> group; **9** exhibits an isotopically sensitive band at 1002  $\text{cm}^{-1}$ , which shifts to 945  $\text{cm}^{-1}$  upon substitution of <sup>16</sup>O with <sup>18</sup>O. The observed O–O stretching frequency of **9** (1002  $\text{cm}^{-1}$ ) is significantly lower than that of **8** (1131  $\text{cm}^{-1}$ ). The X-ray crystal structure of **9** revealed the mononuclear side-on Ni(III)–peroxo complex in a distorted octahedral geometry (Figure 1G).<sup>35</sup> The O–O bond length of **9** (1.386(4)  $\text{\AA}$ ) is longer than those of Ni(II)–superoxo complexes, such as **8** (1.301  $\text{\AA}$ , obtained from DFT calculations)<sup>34</sup> and NiO<sub>2</sub> with a  $\beta$ -diketiminato ligand (1.347  $\text{\AA}$ ).<sup>43</sup> Further, the Ni–O average bond distance of **9** (1.889  $\text{\AA}$ ) is shorter than that of **8** (1.984  $\text{\AA}$ , obtained from DFT calculations),<sup>34</sup> supporting a Ni<sup>3+</sup> oxidation state, which was also shown by XAS experiments and supported by DFT calculations. All four *N*-methyl groups of the 12-TMC ligand are oriented *syn* to

**SCHEME 4.** Preferred Geometry of M–O<sub>2</sub> Complex Depending on the Oxidation State of Metal Ion

the peroxo ligand in **9**. There is a correlation between the observed O–O bond length and the O–O stretching frequency in **9**, as observed in other side-on M–O<sub>2</sub> complexes (Figure 2).<sup>12</sup>

**2.1. Metal Ion Effect.** We have shown above that the nature of O<sub>2</sub> binding (e.g., side-on vs end-on and superoxo vs peroxo), the O–O and M–O bond distances in the M–O<sub>2</sub> unit, and the orientation of methyl groups in TMC ligands are markedly influenced by the central metal ions of M–O<sub>2</sub> complexes (Figure 1; Tables 1 and 2). For example, in the series of M–O<sub>2</sub> complexes bearing 14-TMC ligand, the O<sub>2</sub> ligands of the Cr(III)–superoxo and Ni(II)–superoxo complexes are bound to the metal ions in an end-on fashion, whereas those of the Mn(III)–peroxo, Fe(III)–peroxo, and Co(III)–peroxo complexes are coordinated to the metal ions in a side-on fashion (see the structures in the first row in Table 1). Thus, there is a trend that metal ions with a relatively high oxidation state prefer to form a side-on metal–peroxo complex, probably due to the stronger donor interaction by the peroxo ligand (e.g., Mn(III)–O<sub>2</sub>, Fe(III)–O<sub>2</sub>, and Co(III)–O<sub>2</sub>), whereas an end-on metal–superoxo complex is a preferred structure for metal ions with a relatively low oxidation state (e.g., Ni(II)–O<sub>2</sub>) (see Scheme 4).

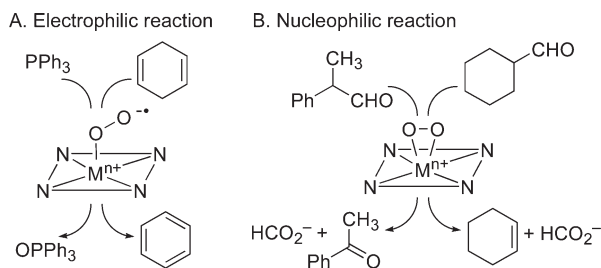
Second, the M–O and O–O bond distances are inversely correlated ( $R^2 = 0.794$ ), although there is one exception with **4** (Figure 3). Such a correlation has been observed in other mononuclear side-on metal–O<sub>2</sub> complexes, and the correlation was interpreted with the orbital overlap between in-plane O<sub>2</sub>  $\pi^*$  orbital and metal  $d_{xy}$  orbital.<sup>12</sup> Thus, the decrease of the O–O bond distances is correlated with the increase of the M–O bond distances, resulting from the decrease of metal–oxygen covalency. Co(III)–peroxo (**7**) and Ni(III)–peroxo (**9**) complexes bearing 12-TMC ligand represent another good example for the correlation between the M–O and O–O bond distances. The average Co–O bond distance of **7** (1.866 Å) is shorter than the average Ni–O bond distance of **9** (1.889 Å), whereas the O–O bond distance of **7** (1.4389(17) Å) is longer than that of **9** (1.386(4) Å).<sup>33,35</sup> A similar trend in M–O bond distances has been observed in Mn(III)–peroxo (**2**) and Fe(III)–peroxo (**4**) complexes with 14-TMC. The Mn–O bond distance of **2**

**FIGURE 3.** Plot of M–O versus O–O bond distances (Å) of M–O<sub>2</sub> complexes. The dashed line represents a least-squares linear fit ( $y = -0.52x + 2.61$ ,  $R^2 = 0.794$ ) of the experimental data except **4**, although the data point of **4** is included in the diagram.

(1.884(2) Å) is shorter than the average Fe–O bond distance of **4** (1.910 Å),<sup>28,31</sup> which is comparable to those of previously reported Mn–O(H) and Fe–O(H) complexes.<sup>44</sup> However, the O–O bond distance of **4** (1.463(6) Å) is significantly longer than that of **2** (1.403(4) Å).<sup>28,31</sup> The extraordinary O–O bond elongation in **4**, which is supported by the rRaman data (vide supra), makes it a dramatic exception to the correlation otherwise demonstrated in Figure 3.

Third, all four *N*-methyl groups of TMC ligands in metal–peroxo complexes point to the same side of the peroxo moiety irrespective of metal ions and the ring size of TMC ligands (see Figure 1).<sup>28,29,31,33,35</sup> In contrast, all *N*-methyl groups of the 14-TMC ligand in the Cr–superoxo complex, [Cr<sup>III</sup>(14-TMC)(O<sub>2</sub>)(Cl)]<sup>+</sup> (**1**), are oriented *anti* to the superoxo ligand.<sup>27</sup> Thus, the orientation of the *N*-methyl groups of TMC ligands in M–O<sub>2</sub> complexes may be determined by the binding mode of the O<sub>2</sub> ligand (e.g., *syn* to the peroxo ligand vs *anti* to the superoxo ligand), although there is only one crystal structure of a M–superoxo species binding a chloride ligand *trans* to the O<sub>2</sub> group (i.e., [Cr<sup>III</sup>(14-TMC)(O<sub>2</sub>)(Cl)]<sup>+</sup>).<sup>27</sup>

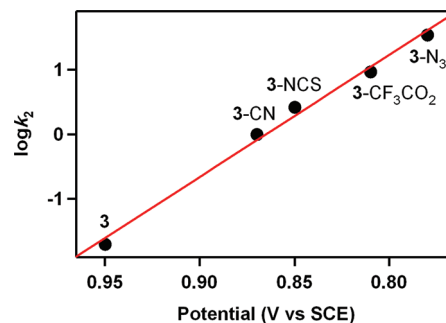
**2.2. Ring Size Effect.** The macrocyclic ligands, *n*-TMC ( $n = 12, 13, 14$ ), give several transition metal–O<sub>2</sub> complexes in different oxidation states, suggesting that the ring size of the macrocycles is an important factor that controls the nature of M–O<sub>2</sub> geometric and electronic structures (e.g., M<sup>*m*+</sup>-superoxo vs M<sup>(*m*+1)+</sup>-peroxo in Scheme 4).<sup>34,35</sup> To understand the ring size effect, XAS studies and DFT calculations were performed on [Co(*n*-TMC)(O<sub>2</sub>)]<sup>+</sup> and [Ni(*n*-TMC)(O<sub>2</sub>)]<sup>+</sup> complexes, where  $n = 12$  or 14.<sup>36</sup> In the case of [Ni(*n*-TMC)(O<sub>2</sub>)]<sup>+</sup> (e.g., **8** and **9**), Ni K-edge energies combined with extended X-ray absorption fine structure (EXAFS) data revealed that the smaller 12-TMC macrocycle forms a Ni(III)–peroxo complex, while the bigger 14-TMC ligand forms a Ni(II)–superoxo species. In **9**, the smaller ring size allows O<sub>2</sub>

**SCHEME 5.** Reactivities of End-on Metal–Superoxo and Side-on Metal–Peroxo Complexes

to bind in a side-on fashion and stabilizes the high oxidation state of  $\text{Ni}^{3+}$ . In the case of  $[\text{Co}(n\text{-TMC})(\text{O}_2)]^+$ , **7** bearing the 12-TMC ligand forms a side-on bound Co(III)–peroxo complex similar to **9** in the nickel system. However, while **8** forms a Ni(II)–superoxo complex, **5** forms a Co(III)–peroxo complex despite having the 14-TMC ligand. This is confirmed by Co K-edge XAS data, which show similar pre-edge and edge energy positions for **5** and **7**.<sup>36</sup> This discrepancy is investigated by DFT calculations, which show that there is a significantly higher thermodynamic driving force for the oxidation of Co(II) to Co(III) compared to Ni(II) to Ni(III).<sup>36</sup> Thus, the results discussed above suggest that there is an interplay of several factors in determining the final stable geometric and electronic structures of  $[\text{M}(n\text{-TMC})(\text{O}_2)]^{m+}$  species, such as the macrocyclic ring size effect and the inherent thermodynamic stability of the metal ion in different oxidation states.

### 3. Reactivity

**3.1. Electrophilic and Nucleophilic Reactions.** Reactivities of the synthetic metal–O<sub>2</sub> complexes have been investigated in electrophilic and nucleophilic reactions in order to understand the chemical properties and mechanisms of the oxidation of organic substrates by metal–superoxo and –peroxo species in enzymatic reactions. The substrates used in electrophilic reactions were  $\text{PPh}_3$  and alkylaromatics with weak C–H bond dissociation energies (BDEs) for oxygen atom transfer and hydrogen atom (H-atom) abstraction reactions, respectively, whereas the nucleophilic character of the metal–O<sub>2</sub> complexes was examined in aldehyde deformylation reactions (Scheme 5). In electrophilic reactions, the nickel(II)–superoxo species,  $[\text{Ni}(\text{14-TMC})(\text{O}_2)]^+$  (**8**), oxidized  $\text{PPh}_3$  to produce  $\text{OPPh}_3$  in quantitative yield (Scheme 5A), and the source of oxygen in the product was confirmed using <sup>18</sup>O-labeled  $[\text{Ni}(\text{14-TMC})(^{18}\text{O}_2)]^+$ .<sup>34,45</sup> The electrophilic character of metal–superoxo species was further investigated in H-atom abstraction reactions by the chromium(III)–superoxo complex,  $[\text{Cr}(\text{14-TMC})(\text{O}_2)(\text{Cl})]^+$  (**1**), using substrates with weak



**FIGURE 4.** Plot of  $\log k_2$  of  $[\text{Mn}(\text{13-TMC})(\text{O}_2)(\text{X})]^+$  (**3-X**;  $\text{X} = \text{N}_3^-$ ,  $\text{CF}_3\text{CO}_2^-$ ,  $\text{NCS}^-$ ,  $\text{CN}^-$ ) at 0 °C against  $E_{p,a}$  values of **3-X**.<sup>29</sup>

C–H bond dissociation energies (BDEs), such as xanthene, dihydroanthracene, and cyclohexadiene.<sup>27,40</sup> In the latter reactions, the order of reaction rates ( $k_2$ ) was xanthene > DHA > CHD, and a linear correlation was observed between the  $\log k_2$  values and the C–H BDE of the substrates. In addition, a kinetic isotope effect (KIE) value of 50 was obtained in the oxidation of DHA by **1**. Such a large KIE value, with the good correlation between  $k_2$  and BDE of the substrates, implicates that the H-atom abstraction is the rate-determining step for the C–H bond activation by the Cr-superoxo complex.

Although metal–peroxo complexes are not reactive in electrophilic reactions, they are capable of conducting nucleophilic reactions, such as the deformylation of aldehydes. For example, the reaction of  $[\text{Co}(\text{14-TMC})(\text{O}_2)]^+$  (**5**) with 2-phenylpropionaldehyde (2-PPA) afforded acetophenone as the major product (Scheme 5B).<sup>32</sup> Similar results were obtained in the reaction of cyclohexanecarboxaldehyde (CCA), and product analysis of the resulting solution revealed the formation of cyclohexene (Scheme 5B).<sup>32</sup> The reactivity of **5** was further investigated using benzaldehydes with a series of electron-donating and -withdrawing substituents at the *para* position of the phenyl group (*para*-Y–Ph–CHO; Y = OMe, Me, F, H, Cl). A positive  $\rho^+$  value of 1.8 in the Hammett plot was obtained, which is consistent with the nucleophilic character of **5** in the oxidation of aldehydes. The reactivity of metal–peroxo complexes was further investigated using primary (1°-CHO), secondary (2°-CHO), and tertiary (3°-CHO) aldehydes, and the observed reactivity order of 1°-CHO > 2°-CHO > 3°-CHO supports the nucleophilic character of the metal–peroxo species.<sup>31,35</sup> Other metal–peroxo complexes,  $[\text{M}^{\text{III}}(n\text{-TMC})(\text{O}_2)]^+$ , showed reactivities in the aldehyde deformylation reactions with positive  $\rho^+$  values in the reactions of *para*-X-substituted benzaldehydes and the reactivity order of 1°-CHO > 2°-CHO > 3°-CHO.<sup>28–33,35</sup> The aldehyde deformylation reactions by metal–peroxo

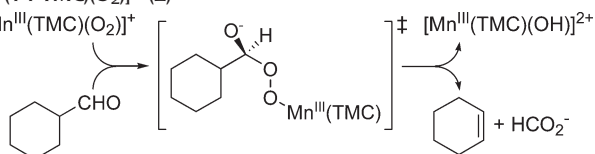
complexes are initiated by attack of the carbonyl group by the M–O<sub>2</sub> moiety (Scheme 6).<sup>14</sup>

A remarkable axial ligand effect on the reactivity of metal–peroxo complexes in heme and nonheme systems was observed in oxidative nucleophilic reactions.<sup>14,29</sup> For example, [Mn(13-TMC)(O<sub>2</sub>)(X)] bearing different axial ligands (**3**-X; X = N<sub>3</sub><sup>-</sup>, CF<sub>3</sub>CO<sub>2</sub><sup>-</sup>, NCS<sup>-</sup>, CN<sup>-</sup>) was studied in the oxidation of CCA and observed the reactivity order of **3**-N<sub>3</sub> > **3**-CF<sub>3</sub>CO<sub>2</sub> > **3**-NCS > **3**-CN > **3**.<sup>29</sup> These results indicate that an increase in the nucleophilicity of the Mn(III)–peroxo unit upon binding of anionic axial ligands makes **3**-X more electron-rich. The electronic effect was confirmed by plotting the potential of **3**-X against the reaction rates of CCA oxidation by **3**-X, showing that the Mn(III)–peroxo complexes with more electron-donating axial ligands are more reactive in oxidative nucleophilic reactions (Figure 4).<sup>29</sup> In addition, the geometric effect on the binding of the anionic ligand *trans* to the peroxo ligand may facilitate the conversion of the side-on peroxo ligand to an end-on peroxo ligand, which probably becomes more nucleophilic.<sup>29</sup> That is, upon binding of an axial ligand, the axial ligand pulls down the metal ion into the TMC ring plane (Scheme 7, A), as proposed in heme models.<sup>14</sup> As a consequence, the peroxo group

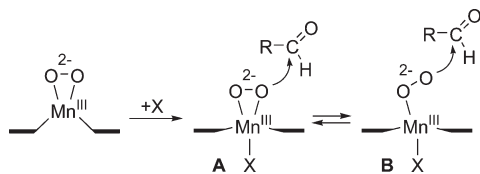
experiences more repulsive interactions from atoms in the TMC ring, resulting in the shift of the equilibrium toward the end-on conformation (Scheme 7, B). DFT calculations on the side-on and end-on Mn(III)–peroxo species illustrate that the peroxo group in the end-on structure has much more anionic character than that in the side-on structure where the electron densities at the oxygen atoms in the side-on and end-on peroxo ligands are –0.24 and –0.63, respectively.<sup>29</sup> Therefore, the oxygen atom in the end-on structure is predicted to be more reactive than the side-on bound structure as a nucleophile.

**3.2. O<sub>2</sub>-Transfer Reactions.** It has been shown that an intermolecular O<sub>2</sub>-transfer occurs from a Ni(III)–peroxo complex, [Ni(12-TMC)(O<sub>2</sub>)]<sup>+</sup> (**9**), to a Mn(II) complex, giving the corresponding Ni(II) complex and Mn(III)–peroxo complex, **2** (Scheme 8).<sup>35</sup> Addition of [Mn(14-TMC)]<sup>2+</sup> to a solution of **9** resulted in the disappearance of the characteristic bands of **9** with the concomitant growth of an absorption band (λ<sub>max</sub> = 453 nm) corresponding to **2** (Figure 5). It was suggested that the O<sub>2</sub>-transfer reaction proceeds via a bimolecular mechanism in which the formation of an undetected [(12-TMC)Ni–O<sub>2</sub>–Mn(14-TMC)]<sup>3+</sup> intermediate is the rate-determining step (Scheme 8).<sup>35</sup>

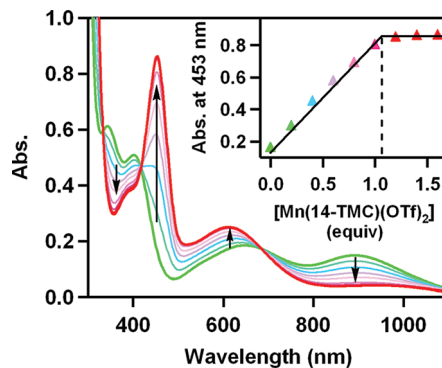
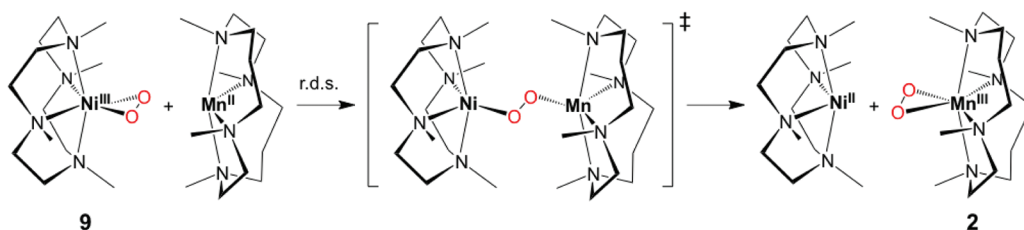
**SCHEME 6.** Proposed Mechanism for the Deformylation of CCA by [Mn<sup>III</sup>(14-TMC)(O<sub>2</sub>)]<sup>+</sup> (**2**)<sup>28</sup>



**SCHEME 7.** Axial Ligand Effect of Metal–Peroxo Complexes in Nucleophilic Reactions<sup>29</sup>



**SCHEME 8.** Reaction Scheme Showing an Intermolecular O<sub>2</sub>-Transfer between Metal Complexes<sup>35</sup>



**FIGURE 5.** UV–vis spectral changes showing the formation of [Mn(14-TMC)(O<sub>2</sub>)]<sup>+</sup> (**2**) (red) and the disappearance of [Ni(12-TMC)(O<sub>2</sub>)]<sup>+</sup> (**9**) (green) by addition of [Mn(14-TMC)]<sup>2+</sup> to a solution of **9**. Inset shows the spectroscopic titration at 453 nm for the formation of **2** as a function of the equivalents of [Mn(14-TMC)]<sup>2+</sup> added to a solution of **9**.<sup>35</sup>



In order to obtain further insight into the O<sub>2</sub>-transfer reaction, isolated [Co(13-TMC)(O<sub>2</sub>)]<sup>+</sup> (**6**) and [Co(12-TMC)(O<sub>2</sub>)]<sup>+</sup> (**7**) complexes were used in the intermolecular O<sub>2</sub>-transfer reactions.<sup>33</sup> The cobalt–peroxo complexes transferred their O<sub>2</sub> unit to [Mn(14-TMC)]<sup>2+</sup>, giving the corresponding Co(II) and Mn(III)–peroxo complexes. Interestingly, the O<sub>2</sub>-transfer from [Co(12-TMC)(O<sub>2</sub>)]<sup>+</sup> (**7**) to [Mn(14-TMC)]<sup>2+</sup> is ~1100 times slower than that from [Ni(12-TMC)(O<sub>2</sub>)]<sup>+</sup> (**9**) to [Mn(14-TMC)]<sup>2+</sup> ( $k_2 = 1.9 \times 10^{-4} \text{ M}^{-1} \text{ s}^{-1}$  for **7** and  $0.2 \text{ M}^{-1} \text{ s}^{-1}$  for **9**), indicating that the Ni(III)–peroxo complex is much more reactive than the Co(III)–peroxo analogue toward the attack of the Mn(II) ion. These results were elucidated by the difference of the enthalpy values of  $65 \text{ kJ mol}^{-1}$  for **7** and  $49 \text{ kJ mol}^{-1}$  for **9**, and the difference of  $\Delta\Delta G$  between **7** and **9** is  $\sim 13 \text{ kJ mol}^{-1}$ . In addition, as the ring size of macrocyclic ligands in the Co(III)–peroxo complexes increases from 12-TMC to 13-TMC, the O<sub>2</sub>-transfer reaction is facilitated, and the reaction rate is increased ~600 times ( $k_2 = 1.9 \times 10^{-4} \text{ M}^{-1} \text{ s}^{-1}$  for **7** and  $1.2 \times 10^{-1} \text{ M}^{-1} \text{ s}^{-1}$  for **6**). This phenomenon in the complete O<sub>2</sub>-transfer between metal complexes bearing TMC ligands is different from the behavior of other systems; in the reactions of metal (M)–O<sub>2</sub> complexes (M = Cu and Ni) and a second metal complex, M or M', formation of homo- or heterodinuclear complexes containing [M<sub>2</sub>( $\mu$ -O)<sub>2</sub>]<sup>n+</sup>, [M<sub>2</sub>(O<sub>2</sub>)]<sup>n+</sup>, or [MM'( $\mu$ -O)<sub>2</sub>]<sup>n+</sup> cores is observed.<sup>37,46</sup>

#### 4. Concluding Remarks

Metal–O<sub>2</sub> intermediates are short-lived and highly reactive in many cases, thus making it difficult to study their chemical and physical properties in the catalytic cycle of dioxygen activation by metalloenzymes. Further, it is especially challenging to obtain crystal structures of the thermally unstable and reactive metal–O<sub>2</sub> intermediates. In this Account, we have described our recent results on the synthesis of a number of metal–O<sub>2</sub> complexes bearing tetraazamacrocyclic TMC ligands and their spectroscopic characterization and crystal structures, by emphasizing the significant role of the macrocyclic TMC ligands in controlling the redox potentials and spin states of metal ions as well as the stability of the bound O<sub>2</sub> ligands. We have also demonstrated the importance of metal ions and the ring size of the supporting ligands on the geometric and electronic structures of the metal–O<sub>2</sub> complexes. Reactivities of the isolated metal–O<sub>2</sub> complexes are discussed in electrophilic and nucleophilic reactions as well as in intermolecular O<sub>2</sub>-transfer reactions. This Account, with the previous review on mononuclear nonheme Fe(IV)–oxo complexes,<sup>47</sup> has shown that the field of nonheme metal–oxygen intermediates (e.g., M–O, M–O<sub>2</sub>, and M–O<sub>2</sub>H)

has been developed greatly in the past decade in bioinorganic chemistry through intense synthetic, spectroscopic, and theoretical efforts; however, there is still much to be explored in trapping, characterizing, and understanding the key intermediates involved in dioxygen activation chemistry by metalloenzymes and their biomimetic compounds.

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#### FOOTNOTES

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#### REFERENCES

- Nam, W. Dioxygen activation by metalloenzymes and models. *Acc. Chem. Res.* **2007**, *40*, 465 and review articles in the special issue.
- Bollinger, J. M., Jr.; Krebs, C. Enzymatic C–H activation by metal-superoxo intermediates. *Curr. Opin. Chem. Biol.* **2007**, *11*, 151–158.
- Kovaleva, E. G.; Lipscomb, J. D. Crystal structures of Fe<sup>2+</sup> dioxygenase superoxo, alkylperoxo, and bound product intermediates. *Science* **2007**, *316*, 453–457.
- Karlsson, A.; Parales, J. V.; Parales, R. E.; Gibson, D. T.; Eklund, H.; Ramaswamy, S. Crystal structure of naphthalene dioxygenase: Side-on binding of dioxygen to iron. *Science* **2003**, *299*, 1039–1042.

- 5 Solomon, E. I.; Brunold, T. C.; Davis, M. I.; Kemsley, J. N.; Lee, S.-K.; Lehnert, N.; Neese, F.; Skulan, A. J.; Yang, Y.-S.; Zhou, J. Geometric and electronic structure/function correlations in non-heme iron enzymes. *Chem. Rev.* **2000**, *100*, 235–349.
- 6 Costas, M.; Mehn, M. P.; Jensen, M. P.; Que, L., Jr. Dioxygen activation at mononuclear nonheme iron active sites: Enzymes, models, and intermediates. *Chem. Rev.* **2004**, *104*, 939–986.
- 7 Mirica, L. M.; Ottenwaelder, X.; Stack, T. D. P. Structure and spectroscopy of copper-dioxygen complexes. *Chem. Rev.* **2004**, *104*, 1013–1045.
- 8 Cramer, C. J.; Tolman, W. B. Mononuclear Cu–O<sub>2</sub> complexes: Geometries, spectroscopic properties, electronic structures, and reactivity. *Acc. Chem. Res.* **2007**, *40*, 601–608.
- 9 Suzuki, M. Ligand effects on dioxygen activation by copper and nickel complexes: Reactivity and intermediates. *Acc. Chem. Res.* **2007**, *40*, 609–617.
- 10 Himes, R. A.; Karlin, K. D. Copper–dioxygen complex mediated C–H bond oxygenation: Relevance for particulate methane monooxygenase (pMMO). *Curr. Opin. Chem. Biol.* **2009**, *13*, 119–131.
- 11 Yao, S.; Driess, M. Lessons from isolable nickel(II) precursor complexes for small molecule activation. *Acc. Chem. Res.* **2012**, *44*, 276–287.
- 12 Cramer, C. J.; Tolman, W. B.; Theopold, K. H.; Rheingold, A. L. Variable character of O–O and M–O bonding in side-on ( $\eta^2$ ) 1:1 metal complexes of O<sub>2</sub>. *Proc. Natl. Acad. Sci. U.S.A.* **2003**, *100*, 3635–3640.
- 13 Bakac, A. Kinetic and mechanistic studies of the reactions of transition metal-activated oxygen with inorganic substrates. *Coord. Chem. Rev.* **2006**, *250*, 2046–2058.
- 14 Wertz, D. L.; Valentine, J. S. Nucleophilicity of iron-peroxo porphyrin complexes. *Struct. Bonding (Berlin)* **2000**, *97*, 37–60.
- 15 Peterson, R. L.; Himes, R. A.; Kotani, H.; Suenobu, T.; Tian, L.; Siegler, M. A.; Solomon, E. I.; Fukuzumi, S.; Karlin, K. D. Cupric superoxo-mediated intermolecular C–H activation chemistry. *J. Am. Chem. Soc.* **2011**, *133*, 1702–1705.
- 16 Kunishita, A.; Kubo, M.; Sugimoto, H.; Ogura, T.; Sato, K.; Takui, T.; Itoh, S. Mononuclear copper(II)-superoxo complexes that mimic the structure and reactivity of the active center of PHM and D $\beta$ M. *J. Am. Chem. Soc.* **2009**, *131*, 2788–2789.
- 17 Maiti, D.; Lee, D.-H.; Gaoutchenova, K.; Würtel, C.; Holthausen, M. C.; Narducci Sarjeant, A. A.; Sundermeyer, J.; Schindler, S.; Karlin, K. D. Reactions of a copper(II) superoxo complex lead to C–H and O–H substrate oxygenation: Modeling copper–monooxygenase C–H hydroxylation. *Angew. Chem., Int. Ed.* **2008**, *47*, 82–85.
- 18 Lee, Y.-M.; Hong, S.; Morimoto, Y.; Shin, W.; Fukuzumi, S.; Nam, W. Dioxygen activation by a nonheme iron(II) complex: Formation of an iron(IV)-oxo complex via C–H activation by a putative iron(III)-superoxo species. *J. Am. Chem. Soc.* **2010**, *132*, 10668–10670.
- 19 Wu, A. J.; Penner-Hahn, J. E.; Pecoraro, V. L. Structural, spectroscopic, and reactivity models for the manganese catalases. *Chem. Rev.* **2004**, *104*, 903–938.
- 20 VanAtta, R. B.; Strouse, C. E.; Hanson, L. K.; Valentine, J. S. Peroxo-(tetraphenylporphinato)manganese(III) and chloro(tetraphenylporphinato)manganese(II) anions. Syntheses, crystal structures, and electronic structures. *J. Am. Chem. Soc.* **1987**, *109*, 1425–1434.
- 21 Kitajima, N.; Komatsuzaki, H.; Hikichi, S.; Osawa, M.; Moro-Oka, Y. A monomeric side-on peroxo manganese(III) complex: Mn(O<sub>2</sub>)(3,5-*i*-Pr<sub>2</sub>pZf)(HB(3,5-*i*-Pr<sub>2</sub>pZ)<sub>3</sub>). *J. Am. Chem. Soc.* **1994**, *116*, 11596–11597.
- 22 Hikichi, S.; Akita, M.; Moro-oka, Y. New aspects of the cobalt-dioxygen complex chemistry opened by hydrotris(pyrazolyl)borate ligands (Tp<sup>R</sup>): Unique properties of Tp<sup>R</sup>Co-dioxygen complexes. *Coord. Chem. Rev.* **2000**, *198*, 61–87.
- 23 Dickman, M. H.; Pope, M. T. Peroxo and superoxo complexes of chromium, molybdenum, and tungsten. *Chem. Rev.* **1994**, *94*, 569–584.
- 24 Tiné, M. R. Cobalt complexes in aqueous solutions as dioxygen carriers. *Coord. Chem. Rev.* **2012**, *256*, 316–327.
- 25 Bailey, C. L.; Drago, R. S. Utilization of O<sub>2</sub> for the specific oxidation of organic substrates with cobalt(II) catalysts. *Coord. Chem. Rev.* **1987**, *79*, 321–332.
- 26 Barefield, E. K. Coordination chemistry of N-tetraalkylated cyclam ligands—A status report. *Coord. Chem. Rev.* **2010**, *254*, 1607–1627.
- 27 Cho, J.; Woo, J.; Nam, W. An “end-on” chromium(III)-superoxo complex: Crystallographic and spectroscopic characterization and reactivity in C–H bond activation of hydrocarbons. *J. Am. Chem. Soc.* **2010**, *132*, 5958–5959.
- 28 Seo, M. S.; Kim, J. Y.; Annaraj, J.; Kim, Y.; Lee, Y.-M.; Kim, S.-J.; Kim, J.; Nam, W. [Mn(tmc)(O<sub>2</sub>)]<sup>+</sup>: A side-on peroxido manganese(III) complex bearing a non-heme ligand. *Angew. Chem., Int. Ed.* **2007**, *46*, 377–380.
- 29 Annaraj, J.; Cho, J.; Lee, Y.-M.; Kim, S. Y.; Latifi, R.; de Visser, S. P.; Nam, W. Structural characterization and remarkable axial ligand effect on the nucleophilic reactivity of a nonheme manganese(III)–peroxo complex. *Angew. Chem., Int. Ed.* **2009**, *48*, 4150–4153.
- 30 Annaraj, J.; Suh, Y.; Seo, M. S.; Kim, S. O.; Nam, W. Mononuclear nonheme ferric-peroxo complex in aldehyde deformylation. *Chem. Commun.* **2005**, 4529–4531.
- 31 Cho, J.; Jeon, S.; Wilson, S. A.; Liu, L. V.; Kang, E. A.; Braymer, J. J.; Lim, M. H.; Hedman, B.; Hodgson, K. O.; Valentine, J. S.; Solomon, E. I.; Nam, W. Structure and reactivity of a mononuclear non-haem iron(III)-peroxo complex. *Nature* **2011**, *478*, 502–505.
- 32 Jo, Y.; Annaraj, J.; Seo, M. S.; Lee, Y.-M.; Kim, S. Y.; Cho, J.; Nam, W. Reactivity of a cobalt(III)-peroxo complex in oxidative nucleophilic reactions. *J. Inorg. Biochem.* **2008**, *102*, 2155–2159.
- 33 Cho, J.; Sarangi, R.; Kang, H. Y.; Lee, J. Y.; Kubo, M.; Ogura, T.; Solomon, E. I.; Nam, W. Synthesis, structural and spectroscopic characterization, and reactivities of mononuclear cobalt(III)-peroxo complexes. *J. Am. Chem. Soc.* **2010**, *132*, 16977–16986.
- 34 Kieber-Emmons, M. T.; Annaraj, J.; Seo, M. S.; Van Heuvelen, K. M.; Toshi, T.; Kitagawa, T.; Brunold, T. C.; Nam, W.; Riordan, C. G. Identification of an “end-on” nickel-superoxo adduct, [Ni(tmc)(O<sub>2</sub>)]<sup>+</sup>. *J. Am. Chem. Soc.* **2006**, *128*, 14230–14231.
- 35 Cho, J.; Sarangi, R.; Annaraj, J.; Kim, S. Y.; Kubo, M.; Ogura, T.; Solomon, E. I.; Nam, W. Geometric and electronic structure and reactivity of a mononuclear “side-on” nickel(III)–peroxo complex. *Nat. Chem.* **2009**, *1*, 568–572.
- 36 Sarangi, R.; Cho, J.; Nam, W.; Solomon, E. I. XAS and DFT investigation of mononuclear cobalt(III) peroxo complexes: Electronic control of the geometric structure in CoO<sub>2</sub> versus NiO<sub>2</sub> systems. *Inorg. Chem.* **2011**, *50*, 614–620.
- 37 Kieber-Emmons, M. T.; Riordan, C. G. Dioxygen activation at monovalent nickel. *Acc. Chem. Res.* **2007**, *40*, 618–625.
- 38 Rohde, J.-U.; In, J.-H.; Lim, M. H.; Brennessel, W. W.; Bukowski, M. R.; Stubna, A.; Münck, E.; Nam, W.; Que, L., Jr. Crystallographic and spectroscopic characterization of a nonheme Fe(IV)=O complex. *Science* **2003**, *229*, 1037–1039.
- 39 Fukuzumi, S.; Morimoto, Y.; Kotani, H.; Naumov, P.; Lee, Y.-M.; Nam, W. Crystal structure of a metal ion-bound oxoiron(IV) complex and implications for biological electron transfer. *Nat. Chem.* **2010**, *2*, 756–759.
- 40 Cho, J.; Woo, J.; Han, J. E.; Kubo, M.; Ogura, T.; Nam, W. Chromium(V)-oxo and chromium(III)-superoxo complexes bearing a macrocyclic TMC ligand in hydrogen atom abstraction reactions. *Chem. Sci.* **2011**, *2*, 2057–2062.
- 41 Qin, K.; Incarvito, C. D.; Rheingold, A. L.; Theopold, K. H. A structurally characterized chromium(III) superoxide complex features “side-on” bonding. *Angew. Chem., Int. Ed.* **2002**, *41*, 2333–2335.
- 42 Li, F.; Meier, K. K.; Cranswick, M. A.; Chakrabarti, M.; Van Heuvelen, K. M.; Münck, E.; Que, L., Jr. Characterization of a high-spin non-heme Fe<sup>III</sup>-OOH intermediate and its quantitative conversion to an Fe<sup>V</sup>=O complex. *J. Am. Chem. Soc.* **2011**, *133*, 7256–7259.
- 43 Yao, S.; Bill, E.; Milsman, C.; Wieghardt, K.; Driess, M. A “side-on” superoxonickel complex [LNi(O<sub>2</sub>)] with a square-planar tetracoordinate nickel(II) center and its conversion into [LNi( $\mu$ -OH)<sub>2</sub>Ni]. *Angew. Chem., Int. Ed.* **2008**, *47*, 7110–7113.
- 44 MacBeth, C. E.; Gupta, R.; Mitchell-Koch, K. R.; Young, V. G., Jr.; Lushington, G. H.; Thompson, W. H.; Hendrich, M. P.; Borovik, A. S. Utilization of hydrogen bonds to stabilize M–O(H) units: Synthesis and properties of monomeric iron and manganese complexes with terminal oxo and hydroxo ligands. *J. Am. Chem. Soc.* **2004**, *126*, 2556–2567.
- 45 Latifi, R.; Tahsini, L.; Kumar, D.; Sastry, G. N.; Nam, W.; de Visser, S. P. Oxidative properties of a nonheme Ni(II)(O<sub>2</sub>) complex: Reactivity patterns for C–H activation, aromatic hydroxylation and heteroatom oxidation. *Chem. Commun.* **2011**, *47*, 10674–10676.
- 46 Aboeella, N. W.; York, J. T.; Reynolds, A. M.; Fujita, K.; Kinsinger, C. R.; Cramer, C. J.; Riordan, C. G.; Tolman, W. B. Mixed metal bis( $\mu$ -oxo) complexes with [CuM( $\mu$ -O)<sub>2</sub>]<sup>II+III</sup> (M = Ni(III) or Pd(II)) cores. *Chem. Commun.* **2004**, 1716–1717.
- 47 Nam, W. High-valent iron(IV)-oxo complexes of heme and non-heme ligands in oxygenation reactions. *Acc. Chem. Res.* **2007**, *40*, 522–531.